



NCERT



CHAPTER WISE TOPIC WISE

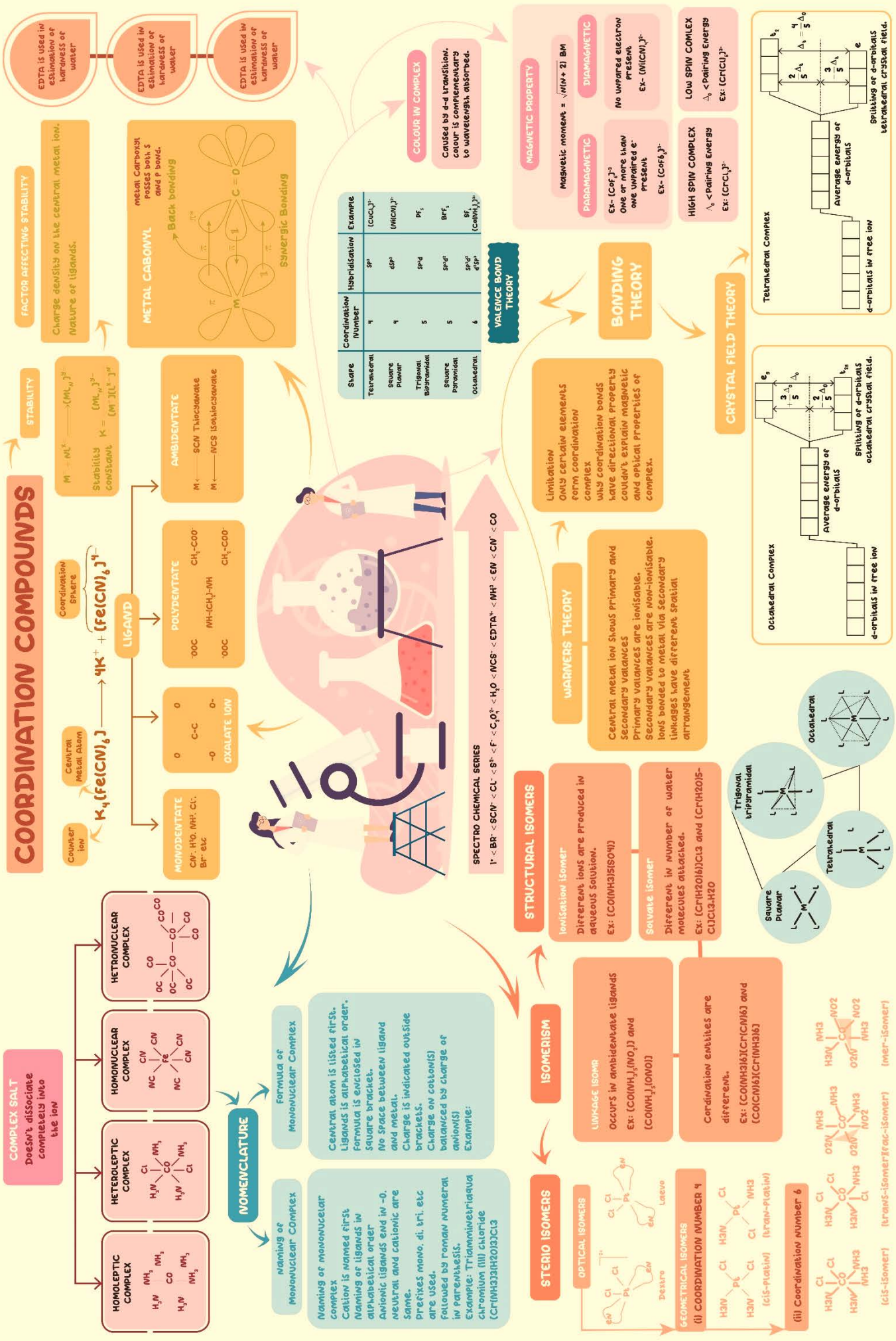
LINE BY LINE QUESTIONS

2024



BY
SCHOOL OF
EDUCATORS

COORDINATION COMPOUNDS



NCERT LINE BY LINE QUESTIONS

(1.) Which of the following coordination compound is also known as heteroleptic complex?

- (a.) $[\text{Co}(\text{NH}_3)_6]^{3+}$ (b.) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]^+$
 (c.) $[\text{PtCl}_4]^{2-}$ (d.) $[\text{Cu}(\text{CN})_4]^{3-}$

(2.) The crystal field stabilisation energy (CFSE) for $[\text{CoClO}]^{4-}$ is 15000cm^{-1} . The CFSE for $[\text{CoCl}_4]^{2-}$ will be

- (a.) 6000cm^{-1} (b.) 16000cm^{-1}
 (c.) 18000cm^{-1} (d.) 8000cm^{-1}

(3.) Match the Column I with Column II and select the correct codes given below.

| Column I | Column II |
|--|---|
| Complex | IUPAC name |
| (P) $[\text{Co}(\text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2)_3]\text{SO}_4$ | (i) Pentaammine carbonatocobalt (III) chloride |
| (Q) $[\text{Co}(\text{NH}_3)_5(\text{CO}_3)]\text{Cl}$ | (ii) Tris(ethane-1,2 -diammine) cobalt (III) sulphate |
| (R) $[\text{CoCl}_2(\text{en})_2]\text{Cl}$ | (iii) Dichloridobis (ethane-1,2-diammine) cobalt (III) chloride |
| (S) $[\text{CoCl}_2(\text{en})_2]^+$ | (iv) Dichloridobis (ethane-1,2-diammine) cobalt (III) |

- (a.) P-(i),Q-(ii),R-(iii),S-(iv) (b.) P-(iv),Q-(iii),R-(ii),S-(i)
 (c.) P-(ii),Q-(i),R-(iii),S-(iv) (d.) P-(iii),Q-(ii),R-(i),S-(iv)

(4.) Due to the presence of ambidentate ligands coordination compounds show isomerism. Palladium complexes of the type $[\text{Pd}(\text{C}_6\text{H}_5)_2(\text{SCN})_2]$ and $[\text{Pd}(\text{C}_6\text{H}_5)_2(\text{NCS})_2]$ are

- (a.) linkage isomers (b.) coordination isomers
 (c.) ionisation isomers (d.) geometrical isomers

(5.) Which of the following formula belongs to Mohr's salt?

- (a.) $\text{KCl} \cdot \text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ (b.) $\text{FeSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$
 (c.) $\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$ (d.) None of these

(6.) Match the column I with column II and select the correct code given below.

| Column I | Column II |
|----------|-----------|
| | |

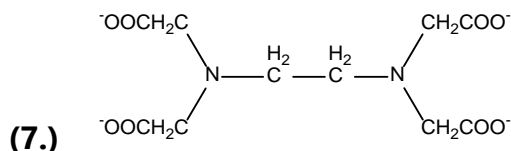
| Hybridisation | Geometry |
|---------------|----------------------------|
| (P) sp^3 | (i) Tetrahedral |
| (Q) dsp^2 | (ii) Square planar |
| (R) sp^3d | (iii) Trigonal bipyramidal |
| (S) sp^3d^2 | (iv) Octahedral |

(a.) P-(i),Q-(ii),R-(iii),S-(iv)

(b.) P-(i),Q-(iii),R-(iv),S-(ii)

(c.) P-(iii),Q-(iv),R-(ii),S-(i)

(d.) P-(ii),□O-(i),R-(iii),S-(iv)



The given compounds is

(a.) hexadentate ligand.

(b.) pentadentate ligand.

(c.) tetradentate ligand.

(d.) didenate ligand.

(8.) What is the oxidation state of the compound tetracarbonyl nickel?

(a.) Zero

(b.) One

(c.) Two

(d.) Four

(9.) Which of the metal carbonyl have metal (M) – metal (M) bond?

(a.) $[\text{Fe}(\text{CO})_5]$

(b.) $[\text{Cr}(\text{CO})_6]$

(c.) $[\text{Ni}(\text{CO})_4]$

(d.) $[\text{Mn}_2(\text{CO})_{10}]$

(10.) Some statements are given below regarding coordination complexes. Select the correct statement.

(I) $[\text{MnCl}_6]^{3-}$ has outer orbital complex.

(II) $[\text{FeF}_6]^{3-}$ involve sp^3d^2 hybridisation.

(III) $[\text{CoF}_6]^{3-}$ has paramagnetic in nature.

(a.) I, II and III

(b.) II and III

(c.) I and III

(d.) I and II

(11.) Which of the following compounds show linkage isomerism

(a.) $[\text{Co}(\text{NH}_3)_5(\text{NO}_2)]\text{Cl}_2$

(b.) $[\text{Co}(\text{NH}_3)_6][\text{Cr}(\text{CN})_6]$

(c.) $[\text{Co}(\text{NH}_3)_5\text{SO}_4]\text{Br}$

(d.) $[\text{Cr}(\text{H}_2\text{O})_6]\text{Cl}_3$

(12.) What are the coordination number of d^2sp^3 and sp^3d^2 hybridised complex respectively?

(a.) 2 and 6

(b.) 4 and 6

(c.) 6 and 6

(d.) 6 and 4

(13.) What is the coordination number of the complex $[\text{Fe}(\text{C}_2\text{O}_4)_3]^{3-}$?

(a.) 4

(b.) 6

(c.) 2

(d.) 0

(14.) IUPAC name of the complex $[\text{Cr}(\text{NH}_3)_3(\text{H}_2\text{O})_3]\text{Cl}_3$ is

(a.) triamminetriaquachromium(II) chloride

(b.) hexaamminerriaquachromium(III) chloride

(c.) diamminetriumminechromium(II) chloride

(d.) triamminetriaquachromium(III) chloride

(15.) 1 mol of $\text{CoCl}_3 \cdot 6\text{NH}_3$ gave how many mol of AgCl precipitate when it is treated with excess silver nitrate solution?

(a.) 3 mol of AgCl

(b.) 2 mol of AgCl

(c.) 1 mol of AgCl

(d.) 4 mol of AgCl

(16.) Match the metal ions given in column I with the spin magnetic moments of the given in column II and assign the correct code:

| Column I | Column II |
|----------------------|-----------------------------|
| (P) Co^{3+} | (i) $\sqrt{8}\text{B.M.}$ |
| (Q) Cr^{3+} | (ii) $\sqrt{35}\text{B.M.}$ |
| (R) Fe^{3+} | (iii) $\sqrt{3}\text{B.M.}$ |
| (S) Ni^{2+} | (iv) $\sqrt{24}\text{B.M.}$ |
| | (v) $\sqrt{15}\text{B.M.}$ |

P Q R S

(a.) (iv) (v) (ii) (i)

(b.) (i) (ii) (iii) (iv)

(c.) (iv) (i) (ii) (iii)

(d.) (iii) (v) (i) (ii)

(17.) Which of the following ligands form a chelate?

(a.) Acetate

(b.) Oxalate

(c.) Cyanide

(d.) Ammonia

(18.) Which of the following statement is correct with respect to $[\text{Ni}(\text{CO})_4]$ and $[\text{NiCl}_4]^{2-}$

(a.) $[\text{Ni}(\text{CO})_4]$ and $[\text{NiCl}_4]^{2-}$ both are square planar

(b.) $[\text{Ni}(\text{CO})_4]$ and $[\text{NiCl}_4]^{2-}$ both are tetrahedral

(c.) $[\text{Ni}(\text{CO})_4]$ is tetrahedral while $[\text{NiCl}_4]^{2-}$ is square planar

(d.) $[\text{Ni}(\text{CO})_4]$ is square planar while $[\text{NiCl}_4]^{2-}$ is tetrahedral

(19.) Which of the following ligands is also considered as ambidentate ligand?

(a.) $\text{C}_2\text{O}_4^{2-}$

(b.) NO_2^-

(c.) NH_3

(d.) H_2O

(20.) Select the incorrect statement

- (a.) Stereoisomerism have the same chemical formula and chemical bonds but they have different spatial arrangement.
- (b.) $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]$ has cis and trans isomerism.
- (c.) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]^+$ has only cis isomerism.
- (d.) Ionisation isomerism is a type of structural isomerism.

(21.) Crystal field theory explains

- (a.) electrostatic model which considers the metal – ligand bond to be ionic.
- (b.) electrostatic model which considers the metal-ligand bond to be covalent.
- (c.) electrostatic model which considers the metal-ligand bond to be ionic and covalent.
- (d.) none of these.

(22.) The correct IUPAC name of $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]$ is

- (a.) diamminedichloridoplatinum (II)
- (b.) diamminedichloridoplatinum (IV)
- (c.) diamminedichloridoplatinum (0)
- (d.) dichloridodiammineplatinum (IV)

(23.) Which of the following arrangement for spectrochemical series is incorrect?

- (a.) $\text{I}^- < \text{Br}^- < \text{SCN}^- < \text{Cl}^-$
- (b.) $\text{S}^{2-} < \text{F}^- < \text{OH}^- < \text{C}_2\text{O}_4^{2-}$
- (c.) $\text{en} < \text{CN}^- < \text{CO} < \text{NCS}^-$
- (d.) $\text{edta}^{4-} < \text{NH}_3 < \text{en} < \text{CN}^-$

(24.) What is the correct IUPAC name of $[\text{Ag}(\text{NH}_3)_2][\text{Ag}(\text{CN})_2]$?

- (a.) Diammine silver(II) dicyanoargentate(II)
- (b.) Diammine silver(I) dicyanoargentate(I)
- (c.) Diammine silver(0) dicyanosilver(II)
- (d.) Diammine silver(II) dicyanoargentate(0)

(25.) Assertion: $[\text{Cu}(\text{H}_2\text{O})_4]^{2+}$ absorb green light.

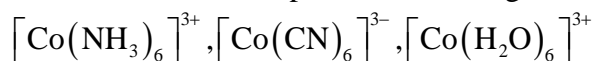
Reason: $[\text{Co}(\text{NH}_3)_5(\text{H}_2\text{O})]^{3+}$ complementary colour is yellow.

- (a.) Both Assertion and Reason are true and Reason is the best explanation of Assertion.
- (b.) Both Assertion and Reason are true but Reason is not the correct explanation of Assertion.
- (c.) Assertion is true but Reason is false.
- (d.) Both Assertion and Reason are false.

(26.) How many ions are produced from the complex $[\text{Co}(\text{NH}_3)_6\text{Cl}_2]$ in solution?

- (a.) 6
- (b.) 4
- (c.) 3
- (d.) 2

(27.) The colour of the coordination compounds depends on the crystal field splitting. What will be the correct order of absorption of wavelength of light in the visible region, for the complexes,



- (a.) $[\text{Co}(\text{CN})_6]^{3-} > [\text{Co}(\text{NH}_3)_6]^{3+} > [\text{Co}(\text{H}_2\text{O})_6]^{3+}$
- (b.) $[\text{Co}(\text{NH}_3)_6]^{3+} > [\text{Co}(\text{H}_2\text{O})_6]^{3+} > [\text{Co}(\text{CN})_6]^{3-}$
- (c.) $[\text{Co}(\text{H}_2\text{O})_6]^{3+} > [\text{Co}(\text{NH}_3)_6]^{3+} > [\text{Co}(\text{CN})_6]^{3-}$
- (d.) $[\text{Co}(\text{NH}_3)_6]^{3+} > [\text{Co}(\text{CN})_6]^{3-} > [\text{Co}(\text{H}_2\text{O})_6]^{3+}$

(28.) Which of the following will not show geometrical isomerism?

- (a.) $[\text{Cr}(\text{NH}_3)_4\text{Cl}_2]\text{Cl}$ (b.) $[\text{Co}(\text{en})_2\text{Cl}_2]\text{Cl}$
(c.) $[\text{Co}(\text{NH}_3)_5\text{NO}_2]\text{Cl}_2$ (d.) $[\text{Pt}(\text{NH}_3)_2\text{C}_2\text{H}_4]\text{Cl}$

(29.) Given below some coordinate complexes with IUPAC name. Identify the incorrect name with their complex.

(I) $[\text{NiCl}_2(\text{PPh}_3)_2]$: Dichloridobis(triphenyl phosphine)nickel(III)

(II) $[\text{Cr}(\text{NH}_3)_3(\text{H}_2\text{O})_3]\text{Cl}_3$: Triamminetriaqua chromium(III) chloride

(III) $\text{Hg}[\text{Co}(\text{SCN})_4]$: Mercury tetrathiocynato cobaltate(II)

- (a.) I and II (b.) I and III
(c.) II and III (d.) I, II and III

(30.) Which one of the following ions exhibits d-d transition and paramagnetism as well?

- (a.) CrO_4^{2-} (b.) $\text{Cr}_2\text{O}_7^{2-}$
(c.) MnO_4^- (d.) MnO_4^{2-}

(31.) Which of the following option is correct with respect to the inner and outer orbital complex?

- (a.) $[\text{Co}(\text{NH}_3)_6]^{3+}$ is inner orbital complex and $[\text{Ni}(\text{NH}_3)_6]^{2+}$ is outer orbital complex.
(b.) $[\text{Ni}(\text{NH}_3)_6]^{2+}$ is inner orbital complex and $[\text{Co}(\text{NH}_3)_6]^{3+}$ is outer orbital complex.
(c.) Both $[\text{Co}(\text{NH}_3)_6]^{3+}$ and $[\text{Ni}(\text{NH}_3)_6]^{2+}$ are inner orbital complex.
(d.) Both $[\text{Co}(\text{NH}_3)_6]^{3+}$ and $[\text{Ni}(\text{NH}_3)_6]^{2+}$ are outer orbital complex.

(32.) The spin only magnetic moment of $[\text{MnBr}_4]^{2-}$ is 5.9 B.M. What is the geometry of the complex?

- (a.) Tetrahedral (b.) Square planar
(c.) Trigonal planar (d.) None of these

(33.) Which isomer in the molecule $[\text{PtCl}_2(\text{en})_2]^{2+}$ shows optical activity?

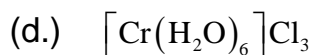
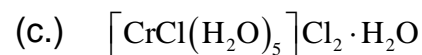
- (a.) Both cis and trans (b.) Neither cis nor trans
(c.) Only cis (d.) Only trans

(34.) Chlorophyll, hemoglobin and vitamin B_{12} are coordination compounds of

- (a.) cobalt, iron and magnesium respectively (b.) magnesium, cobalt and iron respectively
(c.) magnesium, iron and cobalt respectively (d.) iron, cobalt and magnesium respectively.

(35.) When 1 mol $\text{CrCl}_3 \cdot 6\text{H}_2\text{O}$ is treated with excess of AgNO_3 , 3 mol of AgCl are obtained. The formula of the complex is

- (a.) $[\text{CrCl}_3(\text{H}_2\text{O})_3] \cdot 3\text{H}_2\text{O}$ (b.) $[\text{CrCl}_2(\text{H}_2\text{O})_4] \cdot 2\text{H}_2\text{O}$



(36.) Which of the following conditions is necessary for t^3e^1 configuration and what is the nature of ligands?

(a.) $\Delta_o < P$, weak field ligand

(b.) $\Delta_o < P$, strong field ligand

(c.) $\Delta_o > P$, weak field ligand

(d.) $\Delta_o > P$, strong field ligand

(37.) **Assertion:** IUPAC name of $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}_2$ is tetraammineaquachloridocobalt(III) chloride.

Reason: Chlorine is the central element in the compound $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}_2$.

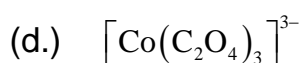
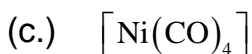
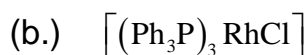
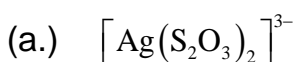
(a.) Both Assertion and Reason are true and Reason is the best explanation of Assertion.

(b.) Both Assertion and Reason are true but Reason is not the correct explanation of Assertion.

(c.) Assertion is true but Reason is false.

(d.) Both Assertion and Reason are false.

(38.) Which of the following complex is also known as Wilkinson catalyst?



(39.) What are the geometry of $[\text{Ni}(\text{CO})_4]$, $[\text{Fe}(\text{CO})_5]$ and $[\text{Cr}(\text{CO})_6]$ respectively?

(a.) Trigonal bipyramidal, tetrahedral, octahedral

(b.) Tetrahedral, trigonal bipyramidal, octahedral

(c.) Octahedral, trigonal bipyramidal, tetrahedral

(d.) Tetrahedral, octahedral, trigonal bipyramidal

(40.) The type of isomerism shown by the complex $[\text{CoCl}_2(\text{en})_2]$ is

(a.) geometrical isomerism

(b.) coordination isomerism

(c.) ionisation isomerism

(d.) linkage isomerism

(41.) $[\text{Co}(\text{NH}_3)_6][\text{Cr}(\text{CN})_6]$, which type of isomerism arises in the given complex?

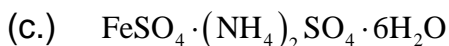
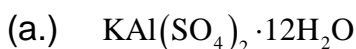
(a.) Linkage isomerism

(b.) Coordination isomerism

(c.) Ionisation isomerism

(d.) Solvate isomerism

(42.) Which of the following is the correct chemical formula of Mohr's salt?



(d.) None of these

(43.) Which of the following ligands are didentate ligand?



(c.) Both (a) and (b)

(d.) Neither (a) nor (b)

(44.) Which of the following statements regarding primary and secondary valences is correct?

(a.) Primary valences are ionisable while secondary valences are non-ionisable.

(b.) Primary valences are non-ionisable while secondary valences are ionisable.

(c.) Both primary and secondary valences are non-ionisable.

(d.) Both primary and secondary valences are ionisable.

(45.) Select the compound which shows ionization isomerism.

- (a.) $[\text{Co}(\text{NH}_3)_5(\text{SCN})]\text{Cl}_2$ (b.) $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{SO}_4$
(c.) $[\text{Co}(\text{NH}_3)_6][\text{Co}(\text{CN})_6]$ (d.) $[\text{Co}(\text{NH}_3)_5(\text{NO}_2)](\text{NO}_3)_2$

(46.) Hexaamminenickel(II) hexanitrocobaltate(III) can be written as

- (a.) $[\text{Ni}(\text{NH}_3)_6]_2[\text{Co}(\text{NO}_2)_6]_3$ (b.) $[\text{Ni}(\text{NH}_3)_6]_3[\text{Co}(\text{NO}_2)_6]_2$
(c.) $[\text{Ni}(\text{NH}_3)_6][\text{Co}(\text{NO}_2)_6]$ (d.) $[\text{Ni}(\text{NH}_3)_6(\text{NO}_2)_6]\text{Co}$

(47.) Out of the following two coordination entities, which is chiral (optically active) .



- (a.) Only I (b.) Only II
(c.) Both I and II (d.) None of these

(48.) Which of the following has longest C-O bond length?

(Free C-O bond length in CO is 1.128 Å)

- (a.) $[\text{Ni}(\text{CO})_4]$ (b.) $[\text{Co}(\text{CO})_4]^-$
(c.) $[\text{Fe}(\text{CO})_4]^{2-}$ (d.) $[\text{Mn}(\text{CO})_6]^+$

(49.) What is the primary valency of the given compounds CrCl_3 , CoCl_2 and PdCl_2 respectively?

- (a.) 2, 2 and 3 (b.) 3, 2 and 2
(c.) 2, 3 and 2 (d.) 3, 2 and 3

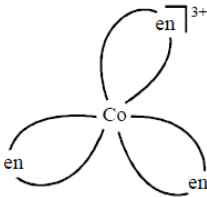
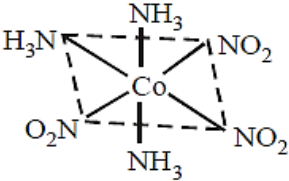
(50.) Select the correct statement

- (a.) The secondary valences are ionisable. (b.) The primary valences are ionisable.
(c.) The primary and secondary valences both are non-ionisable. (d.) None of these.

TOPIC WISE PRACTICE QUESTIONS

TOPIC 1: Coordination Number, Nomenclature and Isomerism

- Select the correct IUPAC name for $[\text{Pt}(\text{C}_5\text{H}_5\text{N})_4]$ both same $[\text{PtCl}_4]$ complex:
1) Tetrapyridineplatinate (II) tetrachloridoplatinate(II)
2) Tetrapyridineplatinum (II) tetrachloridoplatinum(II)
3) Tetrapyridineplatinate (II) tetrachloridoplatinum(II)
4) Tetrapyridineplatinum (II) tetrachloridoplatinate(II)
- $[\text{Pt}(\text{NH}_3)_4\text{Cl}_2] \text{Br}_2$ complex can show :
1) Hydrated as well as ionization isomerism 2) Ionization as well as geometrical isomerism
3) Linkage as well as geometrical isomerism 4) Ionization as well as optical isomerism
- The complex ion $[\text{Pt}(\text{NO}_2)(\text{Py})(\text{NH}_3)(\text{NH}_2\text{OH})]^+$ will give

- 1) 2 isomers (Geometrical) 2) 3 isomers (Geometrical)
 3) 6 isomers (Geometrical) 4) 4 isomers (Geometrical)
4. The geometry of $\text{Ni}(\text{CO})_4$ and $\text{Ni}(\text{PPh}_3)_2\text{Cl}_2$ are
 1) both square planar 2) tetrahedral and square planar
 3) both tetrahedral 4) None of these
5. In the reactions
 $[\text{CoCl}_2(\text{NH}_3)_4]^+ + \text{Cl}^- \rightarrow [\text{CoCl}_3(\text{NH}_3)_3] + \text{NH}_3$ two isomers of the product are obtained. The initial complex is
 1) *cis* isomer 2) *trans* isomer 3) *cis* or *trans* isomers 4) None of these
6. Possible isomerism in complexes $[\text{Co}(\text{NH}_3)_3(\text{NO}_2)_3]$ and $[\text{Co}(\text{NH}_3)_5(\text{NO}_2)]\text{Cl}_2$, respectively are:
 1) Linkage and optical 2) Geometrical and linkage
 3) Optical and ionization 4) Linkage and geometrical
7. Both geometrical and optical isomerisms are shown by
 1) $[\text{Co}(\text{en})_2\text{Cl}_2]^+$ 2) $[\text{Co}(\text{NH}_3)_5\text{Cl}]^{2+}$ 3) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]^+$ 4) $[\text{Cr}(\text{ox})_3]^{3-}$
8. The complex given is

 (i) non-superimposable on its mirror images (ii) optically inactive
 (iii) rotate plane polarised light (iv) planar
 1) (i) and (ii) 2) (i) and (iv) 3) (i), (ii) and (iii) 4) (ii) only
9. An example of double salt is
 1) Bleaching powder 2) $\text{K}_4[\text{Fe}(\text{CN})_6]$ 3) Hypo 4) Potash alum
10. Which of the following type of isomerism is shown by given complex compound?

 1) Facial 2) Meridional 3) Cis 4) Both b and c
11. The tetrahedral complex $[\text{M}(\text{A})_2(\text{X})(\text{Y})]$, where A,B,X and Y are different ligands and M is a metal ion is
 1) optically inactive 2) rotate plane polarized light
 3) incomplete information 4) can't be said
12. Which of the following will give maximum number of isomers?
 1) $[\text{Ni}(\text{C}_2\text{O}_4)(\text{en})_2]^{2-}$ 2) $[\text{Ni}(\text{en})(\text{NH}_3)_4]^{2+}$ 3) $[\text{Cr}(\text{SCN})_2(\text{NH}_3)_4]^+$ 4) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]$
13. Which of the following compounds shows optical isomerism?
 1) $[\text{Co}(\text{CN})_6]^{3-}$ 2) $[\text{Cr}(\text{C}_2\text{O}_4)_3]^{3-}$ 3) $[\text{ZnCl}_4]^{2-}$ 4) $[\text{Cu}(\text{NH}_3)_4]^{2+}$
14. The ionisation isomer of $[\text{Cr}(\text{H}_2\text{O})_4\text{Cl}(\text{NO}_2)]\text{Cl}$ is
 1) $[\text{Cr}(\text{H}_2\text{O})_4(\text{O}_2\text{N})]\text{Cl}_2$ 2) $[\text{Cr}(\text{H}_2\text{O})_4\text{Cl}_2](\text{NO}_2)$
 3) $[\text{Cr}(\text{H}_2\text{O})_4\text{Cl}(\text{ONO})]\text{Cl}$ 4) $[\text{Cr}(\text{H}_2\text{O})_4\text{Cl}_2(\text{NO}_2)] \cdot \text{H}_2\text{O}$
15. A bidentate ligand always
 1) has bonds formed to two metals ions 2) has a charge of +2 or -2
 3) forms complex ions with a charge of +2 or -2

- 4) has two donor atoms forming simultaneously two sigma (s) bonds.
16. Which one of the following complexes is not expected to exhibit isomerism?
 1) $[\text{Ni}(\text{en})_3]^{2+}$ 2) $[\text{Ni}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{2+}$ 3) $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]$ 4) $[\text{Ni}(\text{NH}_3)_2\text{Cl}_2]$
17. The complexes $[\text{Co}(\text{NH}_3)_6]$ $[\text{Cr}(\text{CN})_6]$ and $[\text{Cr}(\text{NH}_3)_6]$ $[\text{Co}(\text{CN})_6]$ are the examples of which type of isomerism?
 1) Linkage isomerism 2) Ionization isomerism
 3) Coordination isomerism 4) Geometrical isomerism
18. For which value of the x, and y, the following square planar compound shows geometrical isomers $[\text{Pt}(\text{Cl})_x(\text{Br})_y]^{2-}$
 1) 1, 3 2) 3, 1 3) 2, 2 4) 1, 1
19. The structure of which of the following chloro species can be explained on the basis of dsp^2 hybridization?
 1) PdCl_4^{2-} 2) FeCl_4^{2-} 3) CoCl_4^{2-} 4) NiCl_4^{2-}
20. Which of the following is not chelating agent?
 1) thiosulphato 2) oxalate 3) glycinate 4) ethylene diamine
21. A similarity between optical and geometrical isomerism is that
 1) each gives equal number of isomers for a given compound
 2) if in a compound one is present then so is the other
 3) both are included in stereoisomerism
 4) they have no similarity
22. Which is the pair of ambidentate ligand?
 1) CN^- , NO_2^- 2) NO_3^- , SCN^- 3) N_3^- , NO_2^- 4) NCS^- , $\text{C}_2\text{O}_4^{2-}$
23. Number of water molecules acting as ligands in $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, $\text{ZnSO}_4 \cdot 5\text{H}_2\text{O}$, $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ respectively are
 1) 5, 5, 7 2) 4, 5, 4 3) 4, 4, 6 4) 4, 4, 7
24. The number of ions formed on dissolving one molecule of $\text{FeSO}_4(\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$ in water is:
 1) 4 2) 5 3) 3 4) 6
25. In octaamine- μ -dihydroxodiiron(III) sulphate, the number of bridging ligands is
 1) 2 2) 1 3) 3 4) None of these

TOPIC 2: Magnetic Moment, Valence Bond Theory and Crystal Field Theory

26. The crystal field splitting energy for octahedral (Δ_o) and tetrahedral (Δ_t) complexes is related as
 1) $\Delta_t = -\frac{1}{2}\Delta_o$ 2) $\Delta_t = -\frac{4}{9}\Delta_o$ 3) $\Delta_t = -\frac{3}{4}\Delta_o$ 4) $\Delta_t = -\frac{2}{5}\Delta_o$
27. Ammonia will not form complex with
 1) Ag^{2+} 2) Pb^{2+} 3) Cu^{2+} 4) Cd^{2+}
28. Which of the following complex compound is low spin, inner orbital, diamagnetic complex?
 1) $[\text{Ni}(\text{NH}_3)_6]\text{Cl}_2$ 2) $\text{K}_3[\text{Fe}(\text{CN})_6]$ 3) $\text{K}_2[\text{PtCl}_6]$ 4) $[\text{Cr}(\text{H}_2\text{O})_6]\text{Cl}_3$
29. As aqueous solution of titanium bromide shows zero magnetic moment. Assuming the complex as octahedral in aqueous solution, the formula of the complex is
 1) $[\text{TiBr}_6]^{3-}$ 2) $[\text{Ti}(\text{H}_2\text{O})_6]\text{Br}_4$ 3) $[\text{TiBr}_6]^{2-}$ 4) $[\text{Ti}(\text{H}_2\text{O})_4\text{Br}_2]$
30. Which of the following option is having maximum number of unpaired electrons –
 1) A tetrahedral d^6 ion 2) $[\text{Co}(\text{NH}_3)_6]^{3+}$
 3) A square planar d^7 ion 4) A co-ordination compound with magnetic moment of 5.92 B.M.
31. The degeneracy of d -orbitals is lost under :
 (I) Strong field ligand (II) Weak field ligand

- (III) Mixed field ligand
 1) I, II and IV 2) I and II (IV) Chelated ligand field
 3) I, II, III and IV 4) I, II and III
32. Relative to the average energy in the spherical crystal field, the t_{2g} orbitals in tetrahedral field is
 1) raised by $(2/5) \Delta_t$ 2) lowered by $(2/5) \Delta_t$
 3) raised by $(3/5) \Delta_t$ 4) lowered by $(1/5) \Delta_t$
33. Which of the following outer orbital complex has highest magnetic moment?
 1) $[\text{Mn}(\text{NH}_3)_6]\text{Cl}_2$ 2) $[\text{Cr}(\text{NH}_3)_6]\text{Cl}_3$ 3) $[\text{Ni}(\text{NH}_3)_6]\text{Cl}_2$ 4) $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$
34. Which of the following are inner orbital complex (i.e., involving d^2sp^3 hybridisation) and is paramagnetic in nature?
 1) $[\text{Mn}(\text{CN})_6]^{3-}$, $[\text{Fe}(\text{CN})_6]^{3-}$, $[\text{Co}(\text{C}_2\text{O}_4)_3]^{3-}$ 2) $[\text{MnCl}_6]^{3-}$, $[\text{FeF}_6]^{3-}$, $[\text{CoF}_6]^{3-}$
 3) $[\text{Mn}(\text{CN})_6]^{3-}$, $[\text{Fe}(\text{CN})_6]^{3-}$ 4) $[\text{MnCl}_6]^{3-}$, $[\text{Fe}(\text{CN})_6]^{3-}$, $[\text{Co}(\text{C}_2\text{O}_4)_3]^{3-}$
35. Mn^{2+} forms a complex with Br^- ion. The magnetic moment of the complex is 5.92 B.M. What would be the probable formula and geometry of the complex?
 1) $[\text{MnBr}_6]^{4-}$, octahedral 2) $[\text{MnBr}_4]^{2-}$, square planar
 3) $[\text{MnBr}_4]^{2-}$, tetrahedral 4) $[\text{MnBr}_5]^{3-}$, trigonal bipyramidal
36. Which of the following hydrate is diamagnetic?
 1) $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ 2) $[\text{Cu}(\text{H}_2\text{O})_6]^{3+}$ 3) $[\text{Co}(\text{NH}_3)_6]^{3+}$ 4) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$
37. Which one of the following will show paramagnetism corresponding to 2 unpaired electrons?
 (Atomic numbers : Ni = 28, Fe = 26)
 1) $[\text{FeF}_6]^{3-}$ 2) $[\text{NiCl}_4]^{2-}$ 3) $[\text{Fe}(\text{CN})_6]^{3-}$ 4) $[\text{Ni}(\text{CN})_4]^{2-}$
38. CN^- is a strong field ligand. This is due to the fact that
 1) it carries negative charge 2) it is a pseudohalide
 3) it can accept electrons from metal species 4) it forms high spin complexes with metal species
39. When pink complex, $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ is dehydrated the colour changes to blue. The correct explanation for the change is :
 1) The octahedral complex becomes square planar. 2) A tetrahedral complex is formed.
 3) Distorted octahedral structure is obtained.
 4) Dehydration results in the formation of polymeric species.
40. The crystal field stabilization energy (CFSE) is the highest for
 1) $[\text{CoF}_4]^{2-}$ 2) $[\text{Co}(\text{NCS})_4]^{2-}$ 3) $[\text{Co}(\text{NH}_3)_6]^{3+}$ 4) $[\text{CoCl}_4]^{2-}$
41. Which of the following complex ion is not expected to absorb visible light ?
 1) $[\text{Ni}(\text{CN})_4]^{2-}$ 2) $[\text{Cr}(\text{NH}_3)_6]^{3+}$ 3) $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$ 4) $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$
42. Crystal field stabilization energy for high spin d^4 octahedral complex is:
 1) $-1.8 \Delta_0$ 2) $-1.6 \Delta_0 + P$ 3) $-1.2 \Delta_0$ 4) $-0.6 \Delta_0$
43. Which of the following is incorrect regarding spectrochemical series?
 1) $\text{NH}_3 > \text{H}_2\text{O}$ 2) $\text{F}^- > \text{C}_2\text{O}_4^{2-}$ 3) $\text{NCS}^- > \text{SCN}^-$ 4) $\text{en} > \text{edta}^{4-}$
44. Of the following complex ions, which is diamagnetic in nature?
 1) $[\text{NiCl}_4]^{2-}$ 2) $[\text{Ni}(\text{CN})_4]^{2-}$ 3) $[\text{CuCl}_4]^{2-}$ 4) $[\text{CoF}_6]^{3-}$
45. The d -electron configurations of Cr^{2+} , Mn^{2+} , Fe^{2+} and Co^{2+} are d^4 , d^5 , d^6 and d^7 , respectively. Which one of the following will exhibit minimum paramagnetic behaviour?
 1) $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ 2) $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$ 3) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ 4) $[\text{Cr}(\text{H}_2\text{O})_6]^{2+}$
 (At. nos. Cr = 24, Mn = 25, Fe = 26, Co = 27)
46. Which of the following complex compounds will exhibit highest paramagnetic behaviour?
 (At. No. : Ti = 22, Cr = 24, Co = 27, Zn = 30)

- 1) $[\text{Ti}(\text{NH}_3)_6]^{3+}$ 2) $[\text{Cr}(\text{NH}_3)_6]^{3+}$ 3) $[\text{Co}(\text{NH}_3)_6]^{3+}$ 4) $[\text{Zn}(\text{NH}_3)_6]^{2+}$
47. Which one of the following is an outer orbital complex and exhibits paramagnetic behaviour ?
 1) $[\text{Ni}(\text{NH}_3)_6]^{2+}$ 2) $[\text{Zn}(\text{NH}_3)_6]^{2+}$ 3) $[\text{Cr}(\text{NH}_3)_6]^{3+}$ 4) $[\text{Co}(\text{NH}_3)_6]^{3+}$
48. Which of the following is the limitation of valence bond theory?
 1) It does not distinguish between weak and strong ligands.
 2) It does not give quantitative interpretation of magnetic data.
 3) It does not explain the colour exhibited by coordination compounds
 4) All of these
49. Low spin complex of d^6 -cation in an octahedral field will have the following energy :
 1) $\frac{-12}{5}\Delta_0 + P$ 2) $\frac{-12}{5}\Delta_0 + 3P$ 3) $\frac{-2}{5}\Delta_0 + 2P$ 4) $\frac{-2}{5}\Delta_0 + P$
 (Δ_0 = Crystal Field Splitting Energy in an octahedral field, P = Electron pairing energy)
50. Among the following complexes the one which shows zero crystal field stabilization energy (CFSE):
 1) $[\text{Mn}(\text{H}_2\text{O})_6]^{3+}$ 2) $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ 3) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ 4) None of these
51. $[\text{Sc}(\text{H}_2\text{O})_6]^{3+}$ ion is
 1) colourless and diamagnetic 2) coloured and octahedral
 3) colourless and paramagnetic 4) coloured and paramagnetic
52. One mole of the complex compound $\text{Co}(\text{NH}_3)_5\text{Cl}_3$, gives 3 moles of ions on dissolution in water. One mole of the same complex reacts with two moles of AgNO_3 solution to yield two moles of AgCl (s). The formula of the complex is
 1) $[\text{Co}(\text{NH}_3)_3\text{Cl}_3] \cdot 2 \text{NH}_3$ 2) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2] \text{Cl} \cdot \text{NH}_3$
 3) $[\text{Co}(\text{NH}_3)_4\text{Cl}] \text{Cl}_2 \cdot \text{NH}_3$ 4) $[\text{Co}(\text{NH}_3)_5\text{Cl}] \text{Cl}_2$
53. How many EDTA (ethylenediaminetetraacetic acid) molecules are required to make an octahedral complex with a Ca^{2+} ion?
 1) One 2) Two 3) Six 4) Three
54. The molar ionic conductances of the octahedral complexes:
 (I) $\text{PtCl}_4 \cdot 5\text{NH}_3$ (II) $\text{PtCl}_4 \cdot 4\text{NH}_3$ (III) $\text{PtCl}_4 \cdot 3\text{NH}_3$ (IV) $\text{PtCl}_4 \cdot 2\text{NH}_3$
 Follow the order
 1) $\text{I} < \text{II} < \text{III} < \text{IV}$ 2) $\text{IV} < \text{III} < \text{II} < \text{I}$ 3) $\text{III} < \text{IV} < \text{II} < \text{I}$ 4) $\text{IV} < \text{III} < \text{I} < \text{II}$
55. The most stable complex among the following is
 1) $\text{K}_3[\text{Al}(\text{C}_2\text{O}_4)_3]$ 2) $[\text{Pt}(\text{en})_2]\text{Cl}$ 3) $[\text{Ag}(\text{NH}_3)_2]\text{Cl}$ 4) $\text{K}_2[\text{Ni}(\text{EDTA})]$
56. Consider the following complex $[\text{Co}(\text{NH}_3)_5\text{CO}_3]\text{ClO}_4$. The coordination number, oxidation number, number of d -electrons and number of unpaired d -electrons on the metal are respectively.
 1) 6, 3, 6, 0 2) 7, 2, 7, 1 3) 7, 1, 6, 4 4) 6, 2, 7, 3
57. Nickel ($Z = 28$) combines with a uninegative monodentate ligand to form a diamagnetic complex $[\text{NiL}_4]^{2-}$. The hybridisation involved and the number of unpaired electrons present in the complex are respectively:
 1) sp^3 , two 2) dsp^2 , zero 3) dsp^2 , one 4) sp^3 , zero

TOPIC 3: Organometallic Compounds

58. The formula of ferrocene is
 1) $[\text{Fe}(\text{CN})_6]^{4-}$ 2) $[\text{Fe}(\text{CN})_6]^{3-}$ 3) $[\text{Fe}(\text{CO})_5]$ 4) $[(\text{C}_5\text{H}_5)_2\text{Fe}]$
59. The organometallic compound is :
 1) $\text{Ti}(\text{OCOCH}_3)_4$ 2) $\text{Ti}(\text{C}_2\text{H}_4)_4$ 3) $\text{Ti}(\text{OC}_6\text{H}_5)_4$ 4) $\text{Ti}(\text{OC}_2\text{H}_5)_4$
60. Which of the following does not have a metal- carbon bond?
 1) $\text{Al}(\text{OC}_2\text{H}_5)_3$ 2) $\text{C}_2\text{H}_5\text{MgBr}$ 3) $\text{K}[\text{Pt}(\text{C}_2\text{H}_4)\text{Cl}_3]$ 4) $\text{Ni}(\text{CO})_4$
61. Which of the following is an organometallic compound?
 1) Lithium methoxide 2) Lithium acetate 3) Lithium dimethylamide 4) Methyl lithium
62. In $\text{Fe}(\text{CO})_5$, the Fe – C bond possesses

- 1) ionic character 2) σ -character only 3) π -character 4) both σ and π characters
63. Which of the following may be considered to be an organometallic compound?
1) Nickel tetracarbonyl 2) Chlorophyll 3) $K_3 [Fe (C_2O_4)_3]$ 4) $[Co (en)_3] Cl_3$
64. $CH_3 - Mg - Br$ is an organometallic compound due to
1) $Mg - Br$ bond 2) $C - Mg$ bond 3) $C - Br$ bond 4) $C - H$ bond.
65. Oxidation state of "V" in $Rb_4K[HV_{10}O_{28}]$ is
1) + 5 2) + 6 3) +7/5 4) + 4
66. For $[Co_2(CO)_8]$, what is the total number of metal – carbon bonds and number of metal–metal bonds.
1) 10, 1 2) 8, 2 3) 8, 1 4) 10, 0
67. Coordination compounds have great importance in biological systems. In this context which of the following statements is incorrect?
1) Cyanocobalamin is B_{12} and contains cobalt
2) Haemoglobin is the red pigment of blood and contains iron
3) Chlorophylls are green pigments in plants and contain calcium
4) Carboxypeptidase - A is an exzyme and contains zinc.
68. An organometallic compound amongst the following is
1) Ferrocene 2) Diethyl zinc 3) Tetraethyl lead (TEL) 4) All of these
69. Which of the following carbonyls will have the strongest C – O bond ?
1) $[Mn (CO)_6]^+$ 2) $[Cr (CO)_6]$ 3) $[V (CO)_6]^-$ 4) $[Fe (CO)_5]$
70. An example of a sigma bonded organometallic compound is :
1) Grignard's reagent 2) Ferrocene 3) Cobaltocene 4) Ruthenocene

NEET PREVIOUS YEARS QUESTIONS

1. Iron carbonyl, $Fe(CO)_5$ is [2018]
1) Tetranuclear 2) Mononuclear 3) Dinuclear 4) Trinuclear
2. The type of isomerism shown by the complex $[CoCl_2(en)_2]$ is [2018]
1) Geometrical isomerism 2) Coordination isomerism
3) Linkage isomerism 4) Ionization isomerism
3. The geometry and magnetic behaviour of the complex $[Ni(CO)_4]$ are [2018]
1) Square planar geometry and diamagnetic 2) Tetrahedral geometry and diamagnetic
3) Tetrahedral geometry and paramagnetic 4) Square planar geometry and paramagnetic
4. An example of a sigma bonded organometallic compound is : [2017]
1) Grignard's reagent 2) Ferrocene 3) Cobaltocene 4) Ruthenocene
5. The correct order of the stoichiometries of $AgCl$ formed when $AgNO_3$ in excess is treated with the complexes : $CoCl_3 \cdot 6NH_3$, $CoCl_3 \cdot 5NH_3$, $CoCl_3 \cdot 4NH_3$ respectively is : [2017]
1) 3 $AgCl$, 1 $AgCl$, 2 $AgCl$ 2) 3 $AgCl$, 2 $AgCl$, 1 $AgCl$
3) 2 $AgCl$, 3 $AgCl$, 1 $AgCl$ 4) 1 $AgCl$, 3 $AgCl$, 2 $AgCl$
6. Correct increasing order for the wavelengths of absorption in the visible region for the complexes of Co^{3+} is : [2017]
1) $[Co(H_2O)_6]^{3+}$, $[Co(en)_3]^{3+}$, $[Co(NH_3)_6]^{3+}$ 2) $[Co(H_2O)_6]^{3+}$, $[Co(NH_3)_6]^{3+}$, $[Co(en)_3]^{3+}$
3) $[Co(NH_3)_6]^{3+}$, $[Co(en)_3]^{3+}$, $[Co(H_2O)_6]^{3+}$ 4) $[Co(en)_3]^{3+}$, $[Co(NH_3)_6]^{3+}$, $[Co(H_2O)_6]^{3+}$
7. Pick out the correct statement with respect to $[Mn(CN)_6]^{3-}$ [2017]
1) It is sp^3d^2 hybridised and tetrahedral 2) It is d^2sp^3 hybridised and octahedral
3) It is dsp^2 hybridised and square planar 4) It is sp^3d^2 hybridised and octahedral
8. Which of the following has longest C–O bond length? (Free C–O bond length in CO is 1.128 Å) [2016]
1) $Ni(CO)_4$ 2) $[Co(CO)_4]^-$ 3) $[Fe(CO)_4]^{2-}$ 4) $[Mn(CO)_6]^+$
9. The sum of coordination number and oxidation number of the metal M in the complex $[M(en)_2(C_2O_4)]Cl$ (where *en* is ethylenediamine) is: [2015]
1) 9 2) 6 3) 7 4) 8
10. The name of complex ion, $[Fe(CN)_6]^{3-}$ is : [2015]
1) Hexacyanoiron (III) ion 2) Hexacyanoferrate (III) ion

c) $[Fe(CN)_6]^{4-}$ iii) 4.90 BM

d) $[Fe(H_2O)_6]^{2+}$ iv) 1.73 BM

23. The IUPAC name of the complex –
[Ag (H₂O)₂] [Ag(CN)₂] is:

[NEET-2022]

1) dicyanidosilver(II) diaquaargentate(II)

2) diaquasilver(II) dicyanidoargentate(II)

3) dicyanidosilver(I) diaquaargentate(I)

4) diaquasilver(I) dicyanidoargentate(I)

24. The order of energy absorbed which is responsible for the color of complexes

[NEET-2022]

A) $[Ni(H_2O)_2(en)_2]^{2+}$ **B)** $[Ni(H_2O)_4(en)]^{2+}$ **and** **c)** $[Ni(en)_3]^{2+}$ **is**

1) A > B > C

2) C > B > A

3) C > A > B

4) B > A > C

NCERT LINE BY LINE QUESTIONS – ANSWERS

| | | | | | | | | | |
|-------|---|-------|---|-------|---|-------|---|-------|---|
| (1.) | b | (2.) | d | (3.) | c | (4.) | a | (5.) | b |
| (6.) | a | (7.) | a | (8.) | a | (9.) | d | (10.) | a |
| (11.) | a | (12.) | c | (13.) | b | (14.) | d | (15.) | a |
| (16.) | a | (17.) | b | (18.) | b | (19.) | b | (20.) | c |
| (21.) | a | (22.) | a | (23.) | c | (24.) | b | (25.) | d |
| (26.) | c | (27.) | c | (28.) | c | (29.) | b | (30.) | d |
| (31.) | a | (32.) | a | (33.) | c | (34.) | c | (35.) | d |
| (36.) | a | (37.) | c | (38.) | b | (39.) | b | (40.) | a |
| (41.) | b | (42.) | c | (43.) | c | (44.) | a | (45.) | b |
| (46.) | b | (47.) | a | (48.) | c | (49.) | b | (50.) | b |

TOPIC WISE PRACTICE QUESTIONS - ANSWERS

| | | | | | | | | | |
|-------|-------|-------|-------|-------|-------|-------|-------|-------|-------|
| 1) 4 | 2) 2 | 3) 2 | 4) 3 | 5) 1 | 6) 2 | 7) 1 | 8) 3 | 9) 4 | 10) 2 |
| 11) 2 | 12) 3 | 13) 2 | 14) 2 | 15) 4 | 16) 4 | 17) 3 | 18) 3 | 19) 1 | 20) 1 |
| 21) 3 | 22) 1 | 23) 3 | 24) 2 | 25) 1 | 26) 2 | 27) 2 | 28) 3 | 29) 2 | 30) 4 |
| 31) 3 | 32) 1 | 33) 1 | 34) 3 | 35) 3 | 36) 3 | 37) 2 | 38) 2 | 39) 2 | 40) 3 |
| 41) 1 | 42) 4 | 43) 1 | 44) 2 | 45) 2 | 46) 3 | 47) 2 | 48) 1 | 49) 4 | 50) 2 |
| 51) 4 | 52) 1 | 53) 4 | 54) 1 | 55) 2 | 56) 4 | 57) 1 | 58) 1 | 59) 4 | 60) 2 |
| 61) 1 | 62) 4 | 63) 4 | 64) 2 | 65) 1 | 66) 1 | 67) 3 | 68) 4 | 69) 1 | 70) 1 |

NEET PREVIOUS YEARS QUESTIONS-ANSWERS

| | | | | | | | | | |
|-------|-------|-------|-------|-------|-------|-------|-------|-------|-------|
| 1) 2 | 2) 1 | 3) 2 | 4) 1 | 5) 2 | 6) 4 | 7) 2 | 8) 3 | 9) 1 | 10) 4 |
| 11) 1 | 12) 3 | 13) 4 | 14) 4 | 15) 2 | 16) 2 | 17) 2 | 18) 4 | 19) 2 | 20) 2 |
| 21) 4 | 22) 3 | 23) 4 | 24) 3 | | | | | | |

NCERT LINE BY LINE QUESTIONS – SOLUTIONS

- (1.) (b) Complexes in which a metal is bound to more than one kind of donor groups, e.g., $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]^+$ are known as heteroleptic. Complexes in which a metal is bound to only one kind of donor groups, e.g., $[\text{Co}(\text{NH}_3)_6]^{3+}$ are known as homoleptic.

- (2.) (d) $\Delta_{\text{tetrahedral}} = \frac{4}{9} \times \Delta_{\text{octahedral}}$
 $= \frac{4}{9} \times 18000 = 8000\text{cm}^{-1}$

- (3.) (c) Correct match is, P-(ii), Q-(i), R-(iii), S-(iv) .

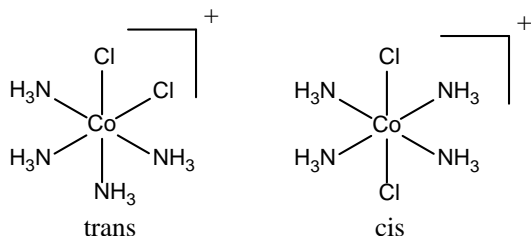
- (4.) (a) Linkage isomers

- (6.) (a)

| Coordination number | Type of hybridisation | Distribution of hybrid orbitals in |
|---------------------|-----------------------|------------------------------------|
|---------------------|-----------------------|------------------------------------|

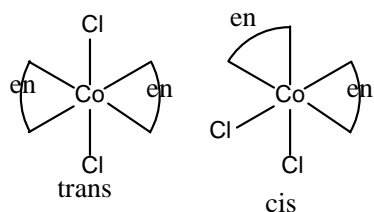
| | | space |
|---|-----------|----------------------|
| 4 | sp^3 | Tetrahedral |
| 4 | dsp^2 | Square planar |
| 5 | sp^3d | Trigonal bipyramidal |
| 6 | sp^3d^2 | Octahedral |
| 6 | d^2sp^3 | Octahedral |

- (7.) (a) The given compound is ethylenediaminetetraacetate ion ($EDTA^{4-}$) is hexadentate ligand.
- (8.) (a) $[Ni(CO)_4]$ – Terracarbonylnickel(0) .
- (10.) (a) $[MnCl_6]^{3-}$, $[FeF_6]^{3-}$ and $[CoF_6]^{3-}$ are outer orbital complexes involving sp^3d^2 hybridisation and paramagnetic corresponding to four, five and four unpaired electrons.
- (11.) (a) $[Co(NH_3)_5(NO_2)]Cl_2$ shows linkage isomerism. $[Co(NH_3)_5(NO_2)]Cl_2$, which is obtained as the red form, in which the nitrite ligands is bound through oxygen ($-ONO$) , and as the yellow form, in which the nitrite ligand is bound through nitrogen ($-NO_2$) .
- (14.) (d) $[Cr(NH_3)_3(H_2O)_3]Cl_3$: Triamminetriaqua chromium(III) chloride.
- (15.) (a) 1mol $CoCl_3 \cdot 6NH_3$ (yellow) gave 3 mol of $AgCl$ when it is treated with excess silver nitrate solution.
- (16.) (a) Co^{3+} , $n = 4$, $\mu = \sqrt{4(4+2)} = \sqrt{24}$ B.M.
 Cr^{3+} , $n = 3$, $\mu = \sqrt{3(3+2)} = \sqrt{15}$ B.M.
 Fe^{3+} , $n = 5$, $\mu = \sqrt{5(5+2)} = \sqrt{35}$ B.M.
 Ni^{3+} , $n = 2$, $\mu = \sqrt{2(2+2)} = \sqrt{8}$ B.M.
- (17.) (b) Among the given ligands oxalate is a bidentate ligand hence forms a chelate.
- (18.) (b) Both $[Ni(CO)_4]$ and $[NiCl_4]^{2-}$ are tetrahedral geometry.
- (20.) (c) $[Co(NH_3)_4Cl_2]^+$ has geometrical isomers, cis and trans.



- (22.) (a) The correct IUPAC name of $[Pr(NH_3)_2Cl_2]$ is diamminedichloridoplatinum(II).
- (23.) (c) In general, ligands can be arranged in a series in the order of increasing field strength as given below:
- $I^- < Br^- < SCN^- < Cl^- < S^{2-} < F^- < OH^- < C_2O_4^{2-} < H_2O < NCS^- < edra^{4-} < NH_3 < en < CN^- < CO$

- (25.) (d) $[\text{Cu}(\text{H}_2\text{O})_4]^{2+}$ absorb red colour and its complementary colour is blue. $[\text{Co}(\text{NH}_3)_5\text{H}_2\text{O}]^{3+}$ absorb blue green colour and its complementary colour is red.
- (27.) (c) The CPSE of the ligands is in the order: $\text{H}_2\text{O} < \text{NH}_3 < \text{CN}^-$ Therefore, excitation energies are in the order:
- $$[\text{Co}(\text{H}_2\text{O})_6]^{3+} < [\text{Co}(\text{NH}_3)_6]^{3+} < [\text{Co}(\text{CN})_6]^{3-}$$
- From the reaction $E = \frac{hc}{\lambda}$
- Hence, $E \propto \frac{1}{\lambda}$
- The order of absorption of wavelength of light in the visible region is:
- $$[\text{Co}(\text{H}_2\text{O})_6]^{3+} > [\text{Co}(\text{NH}_3)_6]^{3+} > [\text{Co}(\text{CN})_6]^{3-}$$
- (28.) (c) Octahedral complex of type $[\text{MA}_5\text{B}_2]$ cannot show geometrical isomerism.
- (29.) (b) $[\text{NiCl}_2(\text{PPh}_3)_2]$: Dichloridobis (triphenyl phosphine) nickel (II)
- $\text{Hg}[\text{Co}(\text{SCN})_4]$: Mercury tetrathiocynato cobaltate(III)
- (30.) (d)
- $\text{CrO}_4^{2-} \text{ Cr}^{6+}$ – Diamagnetic
- $\text{CrO}_7^{2-} \text{ Cr}^{6+}$ – Diamagnetic
- $\text{MnO}_4^- \text{ Mn}^{7+}$ – Diamagnetic
- $\text{MnO}_4^{2-} \text{ Mn}^{6+}$ – Paramagnetic
- In MnO_4^{2-} unpaired electron is present, therefore d-d transition is possible.
- (32.) (a) Since the coordination number of Mn^{2+} ion in the complex ion is 4, it will be either tetrahedral (sp^3 hybridisation) or square planar (dsp^2 hybridisation). But the fact that the magnetic moment of the complex ion is 5.9 B.M., it should be tetrahedral in shape rather than square planar because of the presence of five unpaired electrons in the d-orbitals.
- (33.) (c) In a coordination entity of the type $[\text{PtCl}_2(\text{en})_2]^{2+}$, only the cis-isomer shows optical activity.
- (35.) (d) As 3 moles of AgCl are obtained when 1 mol of $\text{CrCl}_3 \cdot 6\text{H}_2\text{O}$ is treated with excess of AgNO_3 which shows that one molecule of the complex gives three chloride ions in solution. Hence, formula of the complex is $[\text{Cr}(\text{H}_2\text{O})_6]\text{Cl}_3$.
- (37.) (c) Cobalt (Co) is the central metal in the complex $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}_2$.
- (38.) (b) Coordination compounds are used as catalysts for many industrial processes. Examples include rhodium complex, $[(\text{Ph}_3\text{P})_3\text{RhCl}]$ a Wilkinson catalyst is used for the hydrogenation of alkenes.
- (39.) (b) Tetracarbonyl nickel(0) is tetrahedral, pentacarbonyl iron (0) is trigonal bipyramidal while hexacarbonyl chromium (0) is octahedral.
- (40.) (a) $[\text{CoCl}_2(\text{en})_2]$ has geometrical isomerism.



- (42.) (c) Mohr's salt: $\text{FeSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$

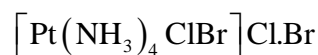
- (44.) (a) The primary valences are normally ionisable and are satisfied by negative ions while secondary valences are non-ionisable. These are satisfied by neutral molecules or negative ions. The secondary valence is equal to the coordination number and is fixed a metal.
- (45.) (b) Ionisation isomerism arises when the counter ion in a complex salt is itself a potential ligand and can displace a ligand which can then become the counter ion. An example is provided by the ionisation isomers $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{SO}_4$ and $[\text{Co}(\text{NH}_3)_5\text{SO}_4]\text{Cl}$.
- (46.) (b) Hexaamminenickel(II) hexanitrocobaltate(III) can be written as $[\text{Ni}(\text{NH}_3)_6]_3[\text{Co}(\text{NO}_2)_6]_2$.
- (47.) (a) Out of $\text{cis}-[\text{CrCl}_2(\text{Ox})_2]^{3-}$ and $\text{trans}-[\text{CrCl}_2(\text{Ox})_2]^{3-}$, cis-isomer is chiral (optically active).
- (48.) (c) $[\text{Fe}(\text{CO})_4]^{2-}$
 Since metal atom is carrying maximum negative charge therefore it would show maximum synergic bonding as a result C–O bond length would be maximum.
- (49.) (b) Primary compounds such as CrCl_3 , CoCl_2 or PdCl_2 have primary valency of 3, 2 and 2 respectively.

TOPIC WISE PRACTICE QUESTIONS - SOLUTIONS

1. 4)

2. 2) $[\text{Pt}(\text{NH}_3)_4\text{Cl}_2]\text{Br}_2$

Ionization isomer



Geometrical isomers

3. 2)

4. 3)

| | $\text{Ni}(\text{CO})_4$ | $\text{Ni}(\text{PPh}_3)_2\text{Cl}_2$ |
|---------------|--------------------------|--|
| O.S. | Ni^0 | Ni^{2+} |
| E.C. | $[\text{Ar}]3d^84s^2$ | $[\text{Ar}]3d^84s^0$ |
| | Pairing of e^- | No pairing of e^- |
| Hybridization | sp^3 (tetrahedral) | sp^3 (tetrahedral) |

5. 1) The initial complex must have two Cl^- ions in cis position as the third Cl^- ion could replace an ammonia cis to both or trans to one giving two isomers. If two Cl^- ions were in the trans positions, the other four positions which could be replaced are equivalent and only one isomer of the product would be obtained.

6. 2) $[\text{Co}(\text{NH}_3)_3(\text{NO}_2)_3]\text{Cl}_2$ complex exhibits geometrical isomerism (G.I.)

Geometrical isomers = 2 (1 cis + 1 trans)

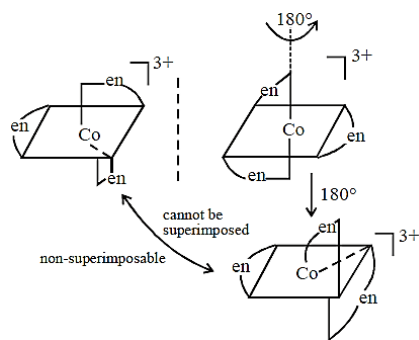
Optical isomers = 0

Space isomers = 2

$[\text{Co}(\text{NH}_3)_5(\text{NO}_2)]\text{Cl}_2$ complex shows linkage and ionization isomerism.

7. 1) The compounds of the type $\text{M}(\text{AA})_2\text{B}_2$ exhibit both geometrical and optical isomerism

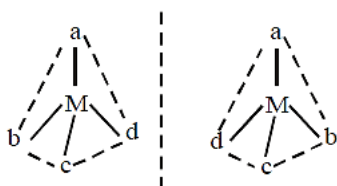
8. (3) Complex is not superimposable on its mirror image hence optically active i.e., rotate plane polarized light.



9. 4)

10. 2) Given compound shows meridional isomerism.

11. (2) Non –superimposable mirror images are optically active, hence rotate plane polarized light.



12. 3) $[\text{Cr}(\text{SCN})_2(\text{NH}_3)_4]^+$ shows linkage, geometrical and optical isomerism. Hence produces maximum no. of isomers.

13. 2)

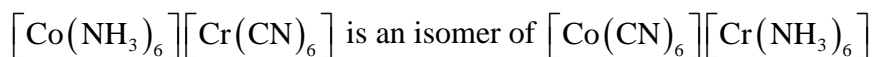
14. 2) Ionisation isomer of $[\text{Cr}(\text{H}_2\text{O})_4\text{Cl}(\text{NO}_2)]\text{Cl}$ is $[\text{Cr}(\text{H}_2\text{O})_4\text{Cl}_2]\text{NO}_2$

15. 4)

16. 4) $[\text{Ni}(\text{NH}_3)_2\text{Cl}_2]$, Ni^{2+} is in sp^3 hybridisation, thus tetrahedral in shape. Hence the four ligands are not different to exhibit optical isomerism. In tetrahedral geometry all the positions are adjacent to one another

\therefore geometrical isomerism also is not possible.

17. (3) Coordination isomerism occurs when cationic and anionic complexes of different metal ions are present in a salt. The two isomers differ in the distribution of ligands in cation and anion e.g.



18. 3) Geometrical isomers of following type of square planar complexes is possible. Ma_2b_2 type, Ma_2bc type and Mabcd type.

19. (1) $[\text{PdCl}_4]^{2-}$ is dsp^2 hybridized and square planar in shape.

20. (1) $\text{S}_2\text{O}_3^{2-}$ is a monodentate ligand whereas other ligands are bidentate.

21. (3) Similarity between optical and geometrical isomerism is that both are included in stereo isomerism.

22. (1)

23. 3) Coordination number of Cu^{2+} , Zn^{2+} and Fe^{2+} are 4, 4 and 6 respectively.

i. $[\text{Cu}(\text{H}_2\text{O})_4]\text{SO}_4 \cdot \text{H}_2\text{O}$; ii. $[\text{Zn}(\text{H}_2\text{O})_4]\text{SO}_4 \cdot \text{H}_2\text{O}$; iii. $[\text{Fe}(\text{H}_2\text{O})_6]\text{SO}_4 \cdot \text{H}_2\text{O}$

24. (2) It is a double salt: $\text{FeSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O} \rightarrow \text{Fe}^{2+} + 2\text{SO}_4^{2-} + 2\text{NH}_4^+$

25. 1)

26. 2)

26. (2) The crystal field splitting in tetrahedral complexes is lower than that in octahedral complexes, and

$$\Delta_t = -\frac{4}{9}\Delta_0$$

27. (2) The complex formation is a characteristic of *d*-block elements. Lead is a *p*-block element hence does not form complex compounds.

28. (3) $[\text{Ni}(\text{NH}_3)_6]\text{Cl}_2$ paramagnetic, outer orbital complex, sp^3d^2

$\text{K}_3[\text{Fe}(\text{CN})_6]$ paramagnetic, inner orbital complex, $d^2 sp^3$

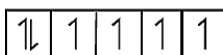
$\text{K}_2[\text{PtCl}_6]$ diamagnetic, inner orbital complex, $d^2 sp^3$

$[\text{Cr}(\text{H}_2\text{O})_6]\text{Cl}_3$ paramagnetic, inner orbital, $d^2 sp^3$

29. (2)

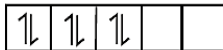
30. (4)

(a) For tetrahedral d^6 ion,



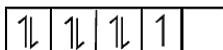
4 unpaired electrons

(b) For $[\text{Co}(\text{NH}_3)_6]^{3+}$,



0 unpaired electrons

(c) For square planar d^7 ion,



1 unpaired electrons

(d) B.M. = $\sqrt{n(n+2)}$, n = unpaired electrons

$$5.92 \text{ B.M.} = \sqrt{n(n+2)}$$

$n = 5$ unpaired electrons

31. 3) Degenerate *d*-orbitals undergo splitting under ligand field created by strong, weak or mixed ligands.

32. 1) t_{2g} orbitals have higher energy by $(2/5)\Delta_t$

33. 1) $[\text{Mn}(\text{NH}_3)_6]\text{Cl}_2$

It is an outer orbital complex ; $\text{Mn}^{2+} (d^5)$

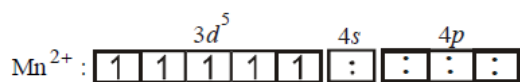
| | | | | |
|------------|------------|------------|------------|------------|
| \uparrow | \uparrow | \uparrow | \uparrow | \uparrow |
|------------|------------|------------|------------|------------|

 sp^3d^2 -hyd.

$$\text{M.M} = \sqrt{35} \text{ B.M}$$

34. 3) $[\text{Mn}(\text{CN})_6]^{3-}$ and $[\text{Fe}(\text{CN})_6]^{3-}$ are inner orbital complexes and paramagnetic while $[\text{Co}(\text{C}_2\text{O}_4)_3]^{3-}$ is diamagnetic in nature

35. 3) Magnetic moment, $\sqrt{n(n+2)} = 5.92 \text{ B.M} \Rightarrow n \text{ (unpaired electrons)} = 5$



sp^3 hybridization (tetrahedral)

36. 3)

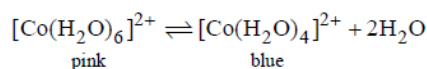
| | | | |
|---------------------------------|---------------------------------|------------------------------|---------------------------------|
| Mn ²⁺ | Cu ³⁺ | Co ³⁺ | Co ²⁺ |
| [Ar]3d ⁵ | [Ar]3d ⁸ | [Ar]3d ⁶ | [Ar]3d ⁷ |
| No pairing of e ⁻ | No pairing of e ⁻ | Pairing of e ⁻ | No pairing of e ⁻ |

Paramagnetic Paramagnetic Diamagnetic Paramagnetic

37. 2) In [NiCl₄]²⁻ chloride ion, being a weak ligand, is not able to pair the electrons in *d* orbital.

38. 2) CN⁻ is a strong field ligand as it is a pseudohalide ion. These ions are strong coordinating ligands and hence have the tendency to form σ-bond (from the pseudo halide to the metal) and π-bond (from the metal to pseudo halide)

39. 2) Hydrated CoCl₂.6H₂O is pink coloured and contains octahedral [Co(H₂O)₆]²⁺ ions. If this is partially dehydrated by heating, then blue coloured tetrahedral ions [Co(H₂O)₄]²⁺ are formed



40. 3) Higher the oxidation state of the metal, greater the crystal field splitting energy. In options 1), 2) and 4), Co is present in + 2 oxidation state and in 3) it is present in + 3 oxidation state and hence has a higher value of CFSE.

41. 1)

42. 4)

*d*⁴ in high spin octahedral complex



$$\text{CFSE} = (-0.4x + 0.6y)\Delta_0$$

Where, $x \rightarrow$ electrons in t_{2g} orbital

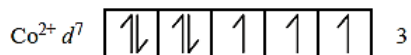
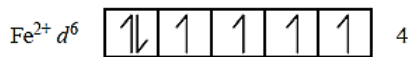
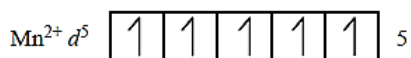
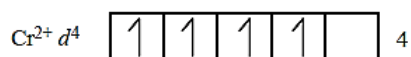
$y \rightarrow$ electrons in e_g orbital

$$\text{CFSE} = [0.6 \times 1] + [-0.4 \times 3] = -0.6 \Delta_0$$

43. 2) According to spectrochemical series $\text{C}_2\text{O}_4^{2-} > \text{F}^-$

44. 2)

45. 3)



Minimum paramagnetic behaviour = [Co (H₂O)₆]²⁺

46. 2)

47. 1) $[\text{Ni}(\text{NH}_3)_6]^{2+}$

$\text{Ni}^{2+} = 3d^8$, according to CFT = $t_2^6 e_g^2$ therefore, hybridisation is sp^3d^2 and complex is paramagnetic.

48. 4)

49. 2) $d^6 : t_{2g}^{2,2,2} e_g^{0,0}$ (in low spin)

$$\text{C.F.S.E} = -0.4 \times 6\Delta_0 + 3P = -\frac{12}{5}\Delta_0 + 3P$$

50. 4) Due to d^5 configuration CFSE is zero

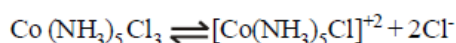
51. 1) $\text{Sc} = [\text{Ar}]3d^1, 4s^2$

Oxidation state of Sc in $[\text{Sc}(\text{H}_2\text{O})_6]^{3+}$ is Sc^{3+}

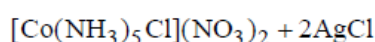
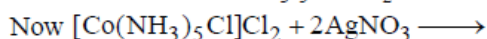
$$\text{Sc}^{3+} = [\text{Ar}]3d^0, 4s^0$$

Thus Sc^{3+} does not have unpaired electron, hence it is diamagnetic and colourless.

52. 4)



\therefore Structure is $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$.



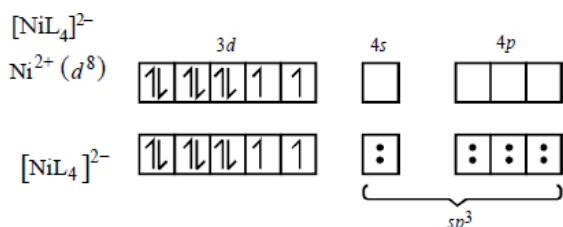
53. 1) EDTA has hexadentate four donor O atoms and 2 donor N atoms and for the formation of octahedral complex one molecule is required

54. 2)

55. 4) $\text{K}_2[\text{Ni}(\text{EDTA})]$ Since EDTA is hexadentate and chelating and coordinates from six sides forming more stable complex.

56. 1) $[\text{Co}(\text{NH}_3)_5\text{CO}_3]\text{ClO}_4$ Six monodentate ligands are attached to Co hence C.N. of Co = 6; O.N. = $x + 5 \times (0) + 1 \times (-2) + 1 \times (-1) = 0 \therefore x = +3$; electronic configuration of $\text{Co}^{3+} [\text{Ar}]3d^6 4s^0$ hence number of d electrons is 6. All d electrons are paired due to strong ligand hence unpaired electron is zero.

57. 1)

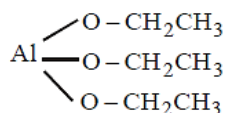


i.e. no. of unpaired electron = 2
hybridization – sp^3 .

58. 4)

59. 2) Organometallic compound is 2) as it contains metalcarbon bonds. In others, direct link of carbon with metal is not present

60. 1) Triethoxyaluminium has no Al – C linkage



61. (4) Compounds that contain at least one carbon metal bond are called organometallic compounds. Hence, based on above definition methyl lithium is an organometallic compound. In other chemical compounds, carbon is not linked with metal.

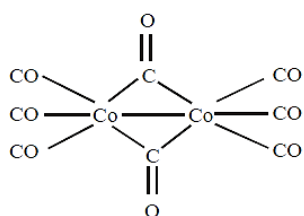
62. (4)

63. (2) Organometallic compounds are those compounds in which a metal is bonded directly to a carbon atom of a molecule. In chlorophyll there is bond between carbon and Mg.

64. (2) Compounds that contain at least one carbon metal bond are known as organometallic compounds. In $\text{CH}_3\text{-Mg-Br}$ (Grignard's reagent) a bond is present between carbon and Mg (Metal) hence it is an organometallic compound.

65. (1)

66. (1) Structure of $[\text{Co}_2(\text{CO})_8]$



67. (3) The chlorophyll molecule plays an important role in photosynthesis, contain porphyrin ring and the metal Mg not Ca.

68. (4)

69. 1) As positive charge on the central metal atom increases, the less readily the metal can donate electron density into the π^* orbitals of CO ligand (donation of electron density into π^* orbitals of CO result in weakening of C – O bond). Hence, the C – O bond would be strongest in $[\text{Mn}(\text{CO})_6]^+$.

70. 1) Grignard's reagent (RMgX) is a σ -bonded organometallic compound.

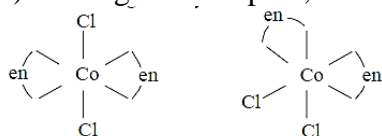
NEET PREVIOUS YEARS QUESTIONS-EXPLANATIONS

1. 2) $\text{Fe}(\text{CO})_5$

$$\text{EAN} = Z - \text{O.N.} + 2(\text{C.N.}) = 26 - 0 + 2(5) = 26 + 10 = 36$$

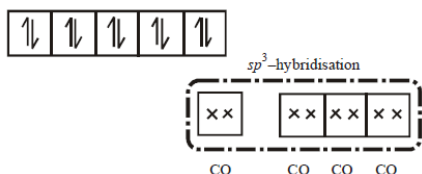
Only one central metal atom/ion is present and it follows EAN rule, so it is mononuclear.

2. 1) In the given complex, the CN of Co is 6, and the complex has octahedral geometry.



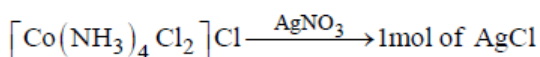
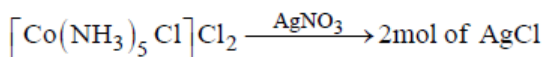
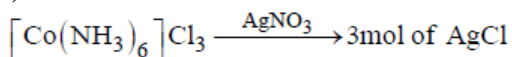
3. 2) $\text{Ni}(28):[\text{Ar}]3d^8 4s^2$

\therefore CO is a strong field ligand, so unpaired electrons get paired. Hence, configuration would be:

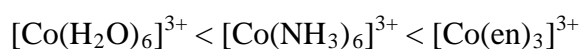


For, four 'CO' ligands hybridisation would be sp^3 and thus the complex would be diamagnetic and of tetrahedral geometry.

4. 1) Grignard's reagent (RMgX) is a s-bonded organometallic compound.
5. 2)



6. 4) The order of the ligand in the spectrochemical series $\text{H}_2\text{O} < \text{NH}_3 < \text{en}$. Hence, the wavelength of the light observed will be in the order



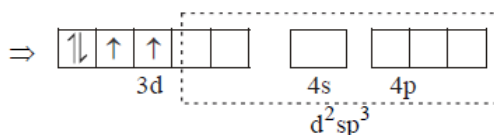
Thus, wavelength absorbed will be in the opposite order



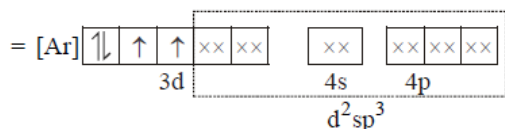
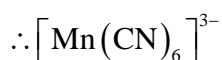
7. 2) In the complex $[\text{Mn}(\text{CN})_6]^{3-}$, O.S. of Mn is +3 E.C. of $\text{Mn}^{+3} \rightarrow 3d^4$



The presence of a strong field ligand CN^- causes pairing of electrons.



As, coordination number of Mn = 6, so it will form an octahedral complex.



8. 3) $[\text{Fe}(\text{CO})_4]^{2-}$ Since metal atom is carrying maximum -ve charge therefore it would show maximum synergic bonding as a resultant C—O bond length would be maximum.

9. $[\text{M}(\text{en})_2(\text{C}_2\text{O}_4)]\text{Cl}$

C_2O_4 = bidentate ligand, carry - 2 charge

en = bidentate ligand, carry 0 charge

\therefore M carry + 3 charge;

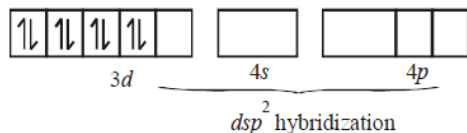
coordination number = 6

$$\therefore \text{Sum} = +3 + 6 = 9$$

10. (4) Carbocation Hexacyanidoerrate (III) ion.

11. 1) $\text{Ni}^{2+} = [\text{Ar}]^{18} 4s^0 3d^8$

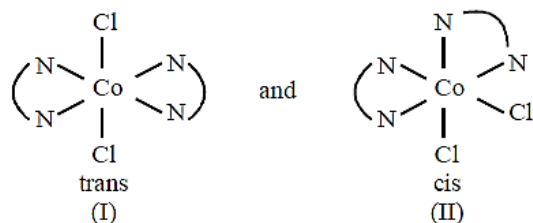
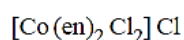
Valence bond theory can be used to predict shape.



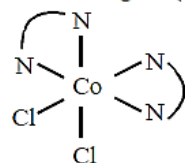
(In presence of ligand, pairing of electron occurs)

\therefore Square planar.

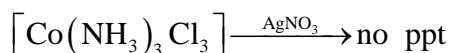
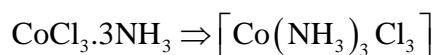
12. 3)



Mirror image of (II)



13. 4) $\text{CoCl}_3 \cdot 3\text{NH}_3$ will not give test for chloride ions with silver nitrate due to absence of ionisable chloride atoms.



14. 4) In $[\text{Co}(\text{CN})_6]^{-3}$ O.N. of Co is +3

$$\therefore \text{Co}^{+3} = 3d^6 4s^0$$

CN^- is a strong field ligand

\therefore Pairing of electrons occurs so in this complex no unpaired electron is present and it is low spin complex.

15. 2) Due to d^5 configuration CFSE is zero.

16. 2)

17. 2)

In $K_4[Fe(CN)_6]$

$Fe(26) = 3d^6 4s^2$

$Fe^{+2} = 3d^6$

in presence of SFL $3d^6 \rightarrow t_{2g}^6 e_g^0$

18. 4)

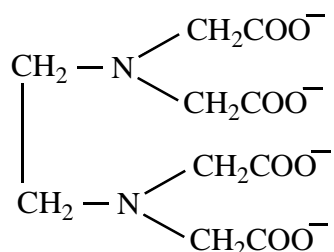
$$\begin{aligned}\Delta_{\text{tetrahedral}} &= \frac{4}{9} \times \Delta_{\text{octahedral}} \\ &= \frac{4}{9} \times 18000 \\ &= 8000 \text{ cm}^{-1}\end{aligned}$$

19. 2) sp^3 - tetrahedral, dsp^2 - square planar

sp^3d - trigonal bipyramidal, d^2sp^3 - octahedral

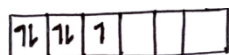
20. 2) According to spectro chemical series . $SCN^- < F^- < C_2O_4^{2-} < CN^-$

21. 4)



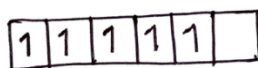
22. 3) $[Fe(CN)_6]^{3-} Fe^{3+} - d^5$

$$S_n = 1 \quad \mu = \sqrt{n(n+2)} = 1.73 \text{ BM}$$



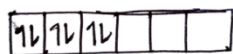
$[Fe(H_2O)_6]^{3+} Fe^{3+} - d^5$

$$n = 5 \quad \mu = \sqrt{n(n+2)} = 5.92 \text{ BM}$$



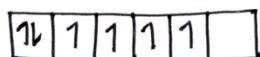
$[Fe(CN)_6]^{4-} Fe^{2+} - d^6$

$$n = 0 \quad \mu = \sqrt{n(n+2)} = 0 \text{ BM}$$



$[Fe(H_2O)_6]^{2+} Fe^{2+} - d^6$

$$n = 4 \quad \mu = \sqrt{n(n+2)} = 4.90 \text{ BM}$$



23. IUPAC name of cpx $[Ag(H_2O)_2][Ag(CN)_2]$ diaquasilver(1) dicyanidoargentate(1)

24. Energy absorbed by complex is directly proportional to strength of ligand

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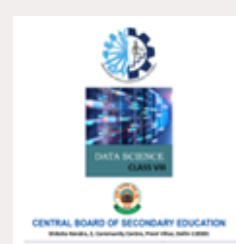
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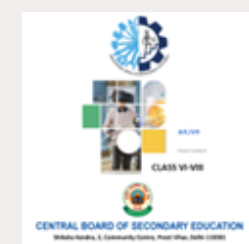
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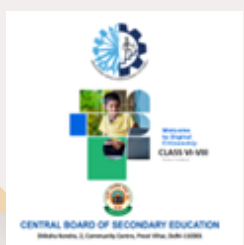
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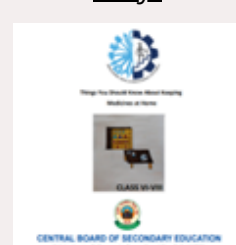
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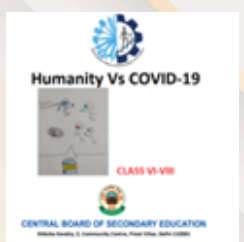
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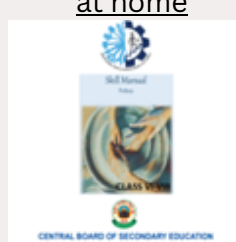
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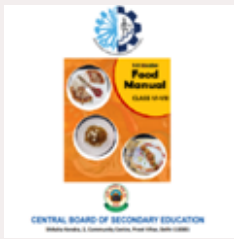
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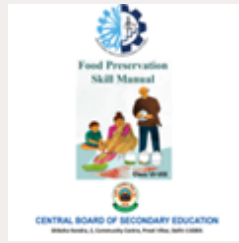
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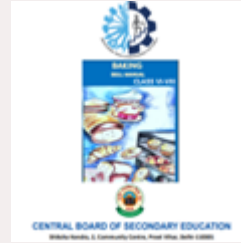
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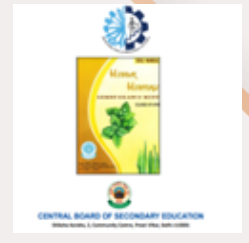
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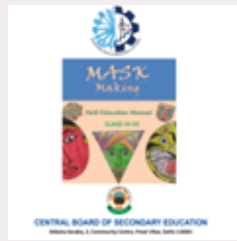
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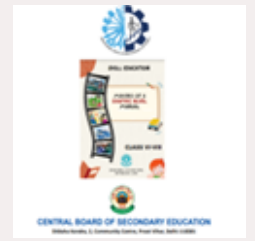
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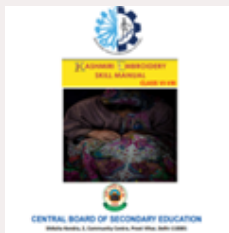
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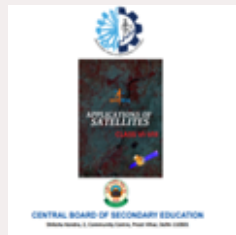
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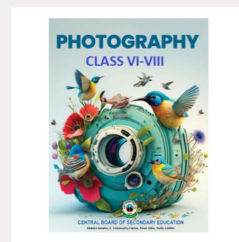
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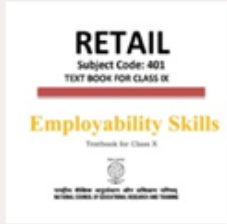


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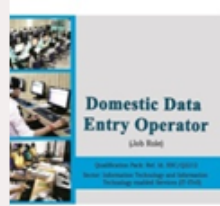


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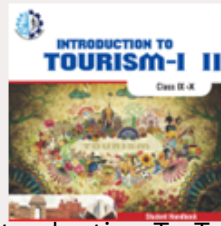
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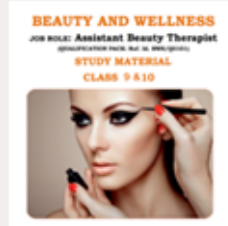
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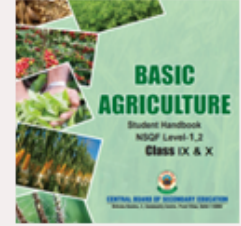
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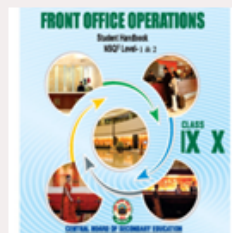
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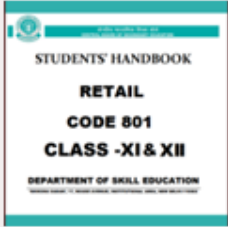


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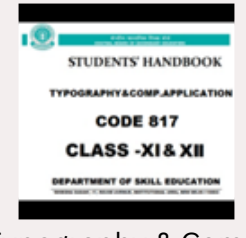
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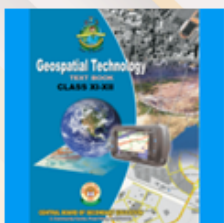
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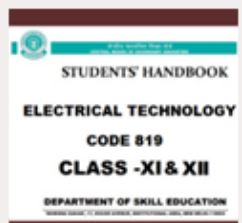
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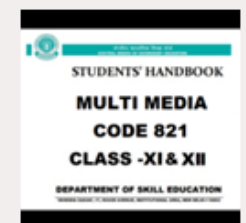
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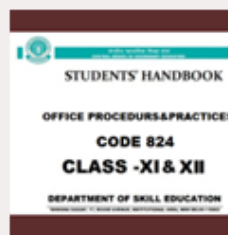
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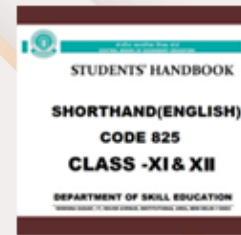
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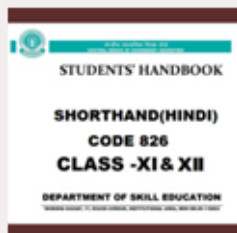
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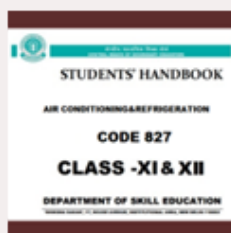
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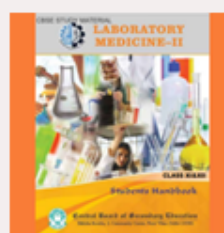
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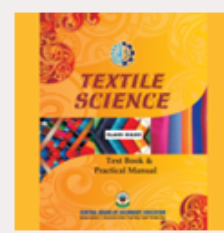
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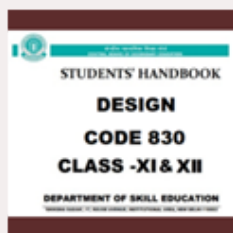
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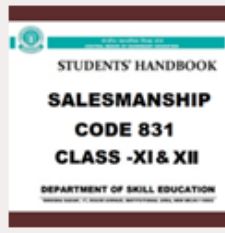
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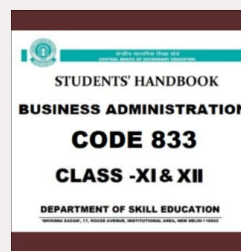
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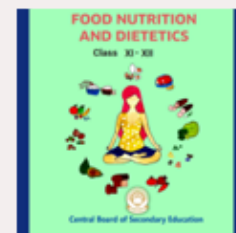
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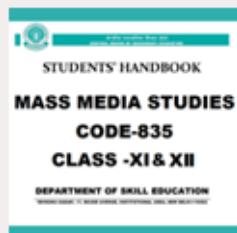
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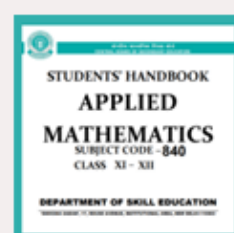
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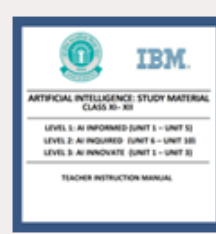
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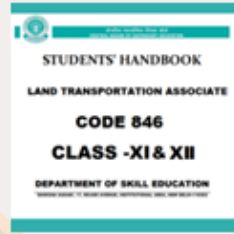
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